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Periodic Table With Chemistry Formulas SparkCharts

* Periodic Table with Chemistry Formulas

sparkcharts

CHEMISTRY FORMULAS

ATOMIC AND MOLECULAR DATA

A. Atomic Number: The number of protons in the nucleus of an atom, also the number of electrons in a neutral atom.

B. Mass Number: The number of protons plus the number of neutrons in the nucleus of an atom.

C. Valence Electrons: The number of electrons in the outermost shell of an atom.

D. Electronegativity: The ability of an atom to attract electrons towards itself.

E. Ionization Energy: The energy required to remove an electron from an atom.

F. Atomic Radius: The distance from the center of the nucleus to the outer edge of the electron cloud.

G. Covalent Radius: The distance between the nuclei of two atoms bonded together, divided by two.

H. Bond Length: The distance between the nuclei of two atoms bonded together.

I. Bond Angle: The angle between two bonds originating from the same central atom.

J. Dipole Moment: A measure of the separation of positive and negative charges in a molecule.

K. Molecular Weight: The sum of the atomic weights of the atoms in a molecule.

L. Empirical Formula: The simplest whole-number ratio of atoms in a compound.

M. Molecular Formula: The actual number of atoms of each element in a molecule.

N. Structural Formula: A diagram showing the arrangement of atoms in a molecule and the bonds between them.

O. Lewis Structure: A diagram showing the valence electrons of atoms in a molecule and how they are shared in bonds.

P. Hybridization: The process of mixing atomic orbitals to form new hybrid orbitals.

Q. Hybridization: The process of mixing atomic orbitals to form new hybrid orbitals.

R. Hybridization: The process of mixing atomic orbitals to form new hybrid orbitals.

S. Hybridization: The process of mixing atomic orbitals to form new hybrid orbitals.

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V. Hybridization: The process of mixing atomic orbitals to form new hybrid orbitals.

W. Hybridization: The process of mixing atomic orbitals to form new hybrid orbitals.

X. Hybridization: The process of mixing atomic orbitals to form new hybrid orbitals.

Y. Hybridization: The process of mixing atomic orbitals to form new hybrid orbitals.

Z. Hybridization: The process of mixing atomic orbitals to form new hybrid orbitals.

1. Atomic Number: Z

2. Mass Number: A

3. Valence Electrons: VE

4. Electronegativity: EN

5. Ionization Energy: IE

6. Atomic Radius: r_a

7. Covalent Radius: r_c

8. Bond Length: d

9. Bond Angle: θ

10. Dipole Moment: μ

11. Molecular Weight: M

12. Empirical Formula: CH_2O

13. Molecular Formula: $C_6H_{12}O_6$

14. Structural Formula: $HO-CH_2-CH_2-OH$

15. Lewis Structure: $H-O-H$

16. Hybridization: sp^3

17. Hybridization: sp^2

18. Hybridization: sp

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99. Hybridization: d^2sp^3

100. Hybridization: d^2sp^3

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Synopsis

SparkCharts®: The information you need—concisely, conveniently, and accurately. Created by Harvard students for students everywhere, these study companions and reference tools cover a wide range of college and graduate school subjects, from Business and Computer Programming to Medicine, Law, and Languages. They'll give you what it takes to find success in school and beyond. Outlines and summaries cover key points, while diagrams and tables make difficult concepts easier to grasp. This two-page chart is a perfect reference for homework and problem sets. On side one, the chart lists the most important chemical formulas and provides quick refreshers on significant figures and balancing equations. Side two includes a beautiful periodic table that gives the following information for each element: Name Atomic number Atomic symbol Atomic mass Oxidation states (most stable state in bold) Electronegativity Electron affinity First ionization potential Atomic radius Electron shell configuration

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Very useful for my chem class at the university of Minnesota

A nice-looking periodic table for quick reference! What more could you want?

Item arrived promptly and as described! Would use again!

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